

Supplementary information for “Data-driven equation for drug–membrane permeability across drugs and membranes”

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S1. DEFINITION OF THE IONIZATION CONSTANTS AND pK_a

Menichetti *et. al.*¹ followed the convention of ChemAxon² while defining apK_a and bpK_a . As it is a bit different from the usual of acidic and basic pK_a s, here we provide the details. The ionization constant K_a and pK_a are defined as

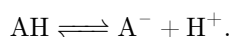
$$K_a = \frac{[\text{conjugate base}] \times [\text{H}^+]}{[\text{conjugate acid}]}, \quad (1)$$

$$pK_a = -\log_{10} K_a = \text{pH} + \log_{10} \frac{[\text{conjugate acid}]}{[\text{conjugate base}]}, \quad (2)$$

where $\text{pH} = -\log_{10}[\text{H}^+]$. In their simulations, Menichetti *et. al.* always started from a neutral compound which, depending on the pH and pK_a , can either protonate or deprotonate. We consider these two cases separately as follows.

S1.1. Deprotonation

A charge-neutral acid AH can deprotonate to release a proton and a charged conjugate base A^- by the following reaction



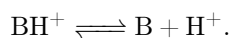
The corresponding acidic pK_a , denoted as apK_a , is defined as

$$K_a = \frac{[\text{A}^-][\text{H}^+]}{[\text{AH}]}, \quad (3)$$

$$apK_a = -\log_{10} K_a = \text{pH} + \log_{10} \frac{[\text{AH}]}{[\text{A}^-]}. \quad (4)$$

S1.2. Protonation

A charge-neutral base B can protonate and becomes a charged conjugate acid $[\text{BH}^+]$ by the following reaction



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To comply with the unified definition of pK_a of ChemAxon—it is the ratio of conjugate acid to conjugate base—the corresponding basic pK_a , denoted as bpK_a , is defined as

$$K_a = \frac{[B][H^+]}{[BH^+]} \quad (5)$$

$$bpK_a = -\log_{10} K_a = pH + \log_{10} \frac{[BH^+]}{[B]} \quad (6)$$

The usefulness of this definition is that now both apK_a and bpK_a are written as $pK_a = pH + \log_{10} \frac{[\text{conjugate acid}]}{[\text{conjugate base}]}$.

Strong acids, as defined in the Acid–base asymptotes section of the paper, have low ($apK_a \leq 4$). At $pH = 7$, from Eq. 4 we find that

$$\frac{[AH]}{[A^-]} = 10^{apK_a-7} \leq 10^{-3}. \quad (7)$$

So, with *decreasing* apK_a , the acid’s concentration $[AH]$ will keep decreasing and the conjugate base’s concentration $[A^-]$ will keep increasing, as expected. Conversely, strong bases have high ($bpK_a \geq 10$). At $pH = 7$, from Eq. 6 we get

$$\frac{[BH^+]}{[B]} = 10^{bpK_a-7} \geq 10^3. \quad (8)$$

So, with *increasing* bpK_a , the base’s concentration $[B]$ will keep decreasing and the conjugate acid’s concentration $[BH^+]$ will keep increasing, again as expected.

S2. DISTRIBUTION OF COMPOUNDS ACROSS PERMEABILITY

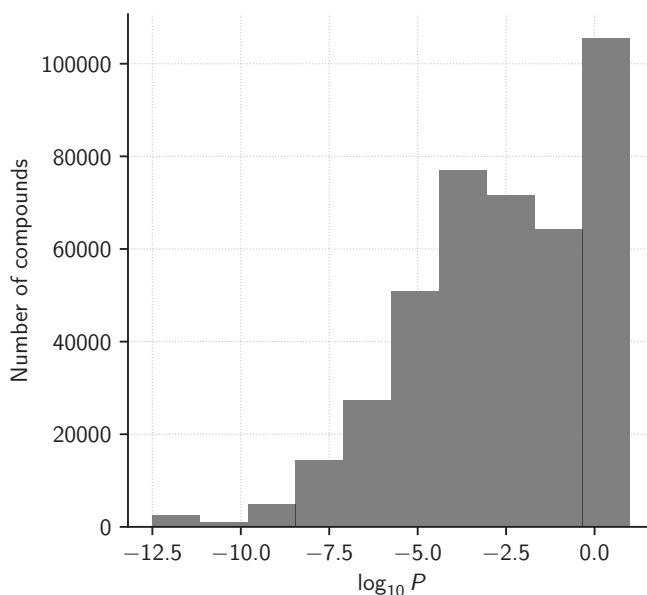


FIG. S1: Distribution of the small molecules considered in this work across the range of permeability values.

S3. TABLE 1 WITH ERROR VALUES

TABLE S1: Table I from the main text along with the standard errors shown in parentheses.

Model	c_0	c_1	c_2	c_3	RMSE	MaxAE	r^2
$f^{\text{Hyd}} = c_0^{\text{Hyd}} + c_1^{\text{Hyd}} \beta \Delta G_{\text{W} \rightarrow \text{O1}}$	-3.444 (± 0.004)	-0.648 (± 0.001)			1.53 (± 0.00)	11.82 (± 0.00)	0.64 (± 0.00)
$f^{1\text{D}} = c_0^{1\text{D}} + c_1^{1\text{D}} (\text{ap}K_{\text{a}} - \text{bp}K_{\text{a}} - 2\beta \Delta G_{\text{W} \rightarrow \text{O1}})$	-5.419 (± 0.007)	0.163 (± 0.000)			1.40 (± 0.00)	6.35 (± 0.02)	0.70 (± 0.00)
$f^{2\text{D}} = c_0^{2\text{D}} + c_1^{2\text{D}} (\sqrt[3]{\beta \Delta G_{\text{W} \rightarrow \text{O1}}} + \beta \Delta G_{\text{W} \rightarrow \text{O1}} - \text{ap}K_{\text{a}})$ $+ c_2^{2\text{D}} (\text{ap}K_{\text{a}}^2 + \text{bp}K_{\text{a}}^2)$	-5.753 (± 0.007)	-0.487 (± 0.001)	-0.017 (± 0.000)		1.06 (± 0.00)	8.28 (± 0.03)	0.83 (± 0.00)
$f^{3\text{D}} = c_0^{3\text{D}} + c_1^{3\text{D}} (\beta \Delta G_{\text{W} \rightarrow \text{O1}} - \text{ap}K_{\text{a}})$ $+ c_2^{3\text{D}} (\text{bp}K_{\text{a}}^2 (\text{ap}K_{\text{a}} + \text{bp}K_{\text{a}}))$ $+ c_3^{3\text{D}} (\text{ap}K_{\text{a}}^2 + (\beta \Delta G_{\text{W} \rightarrow \text{O1}})^2)$	-7.101 (± 0.007)	-0.614 (± 0.002)	-0.001 (± 0.000)	-0.018 (± 0.000)	0.94 (± 0.00)	8.19 (± 0.02)	0.86 (± 0.00)

S4. ONE-DIMENSIONAL DESCRIPTORS

TABLE S2: Best one-dimensional descriptors for ten training sets as predicted by SISSO. Each column corresponds to a particular training set. The data demonstrates the robustness of the predictions across training sets—only twelve unique descriptors are present the best ten descriptors from all training sets. The top three descriptors do not change. The best one-dimensional descriptor ($(\text{ap}K_{\text{a}} - \beta \Delta G_{\text{W} \rightarrow \text{O1}}) - (\text{bp}K_{\text{a}} + \beta \Delta G_{\text{W} \rightarrow \text{O1}})$) and the baseline hydrophobicity descriptor have been highlighted for reference.

1D descriptor	Rank in training set number									
	1	2	3	4	5	6	7	8	9	10
$((\text{ap}K_{\text{a}} - \beta \Delta G_{\text{W} \rightarrow \text{O1}}) - (\text{bp}K_{\text{a}} + \beta \Delta G_{\text{W} \rightarrow \text{O1}}))$	1	1	1	1	1	1	1	1	1	1
$((\text{bp}K_{\text{a}})^2 + (\text{ap}K_{\text{a}} \cdot \beta \Delta G_{\text{W} \rightarrow \text{O1}}))$	2	2	2	2	2	2	2	2	2	2
$(\sqrt[3]{(\text{bp}K_{\text{a}})} + \beta \Delta G_{\text{W} \rightarrow \text{O1}})$	3	3	3	3	3	3	3	3	3	3
$((\text{ap}K_{\text{a}} - \beta \Delta G_{\text{W} \rightarrow \text{O1}}) - \text{bp}K_{\text{a}})$	4	5	4	6	4	4	5	6	5	6
$(\beta \Delta G_{\text{W} \rightarrow \text{O1}} + (\text{bp}K_{\text{a}} + \beta \Delta G_{\text{W} \rightarrow \text{O1}}))$	5	4	6	7	5	5	4	4	4	7
$(\sqrt[3]{(\text{ap}K_{\text{a}})} - \beta \Delta G_{\text{W} \rightarrow \text{O1}})$	6	7	7	5	7	7	7	7	7	5
$(\beta \Delta G_{\text{W} \rightarrow \text{O1}} + \sqrt[3]{(\beta \Delta G_{\text{W} \rightarrow \text{O1}})})$	7	6	5	4	6	6	6	5	6	4
$(\sqrt[3]{(\beta \Delta G_{\text{W} \rightarrow \text{O1}})} - \sqrt[3]{(\text{ap}K_{\text{a}})})$	8	8	8	8	8	8	8	8	8	8
$((\text{ap}K_{\text{a}})^{-1} + \beta \Delta G_{\text{W} \rightarrow \text{O1}})$	9	-	10	-	10	-	-	-	-	-
$(\beta \Delta G_{\text{W} \rightarrow \text{O1}})$	10	9	9	9	9	9	9	9	9	9
$(\sqrt[3]{(\beta \Delta G_{\text{W} \rightarrow \text{O1}})} + (\text{bp}K_{\text{a}} + \beta \Delta G_{\text{W} \rightarrow \text{O1}}))$	-	10	-	-	-	10	-	10	10	10
$((\text{ap}K_{\text{a}})^{-1} - \beta \Delta G_{\text{W} \rightarrow \text{O1}})$	-	-	-	10	-	-	10	-	-	-

